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**Batteries**  
How do battery cells work?

**Why?**  
When several galvanic devices like Zn/Cu galvanic cell galvanic cell connect several devices or provide a flow of electrons. Batteries are the common solution to this challenge. In a battery or voltaic cell, oxidation and reduction reactions provide electrons which power our devices. In this activity we will explore the components of voltaic cells or batteries.

**Model 1 - Voltaic Cell**

1. Consider the reaction in Model 1. Write down the two half-reactions (Ox and Red) in the table on the left before the wire is connected. Explain why the number of electrons (e<sup>-</sup>) must be the same in the two half-reactions.

2. Repeat the same in Model 1 but before the wire is connected and after it is connected. Identify the specific phases of evidence that a chemical reaction has occurred at one point with the wire connected.

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