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K.C.S.E. YEAR 2000
PAPER 2

1. (a) Which one of the following compounds, zinc, ammonia, sugar and copper (II) chloride will conduct an electric current when dissolved in water? Give reasons. (2 marks)

(b) The diagram below shows an electrochemical cell.

Study it and answer the questions that follow.

Given the following

$$\text{Fe}^{3+}/\text{Fe}^{2+} \quad E^\circ = +0.77 \text{ V} \quad \text{Zn}^{2+}/\text{Zn} \quad E^\circ = -0.76 \text{ V}$$

(i) Show on the diagram using arrows, the direction of flow of electrons. (1 mark)

(ii) Name the substances that are used to fill the salt bridge L. (2 marks)

(c) In an experiment to electrolyse zinc sulfate solution with inert electrodes of 0.5 ampere was passed through a solution for 30 minutes. (2 marks)

(i) Give two reasons why it is necessary to electrolyse zinc sulfate solution with silver. (2 marks)

(ii) Calculate the mass of silver that was deposited on the Ag electrode. (Atomic weight of Ag = 108, Faraday constant = 96500 coulombs) (3 marks)

2. Give the name of the following compounds:

(i)

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